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PROCESS FOR TREATING FLUORINE COMPOUND-CONTAINING GAS

INSP

BACKGROUND OF THE INVENTION

1) Field of the Invention

The present invention relates to a process for efficient decomposition treatment of a gas containing
5 fluorine compounds such as C_2F_6 , CF_4 , C_3F_8 , C_4F_8 , CHF_3 , SF_6 , NF_3 , etc. at a low temperature.

2) Related Art

Fluorine compound gases such as CF_4 , C_2F_6 , etc. are used in a large amount as a semiconductor etchant, a
10 semiconductor cleaner, etc. However, it was found that these compounds, once discharged into the atmosphere, turn into ~~warming~~ substances causing global warming. Post-treatment of these compounds after their use ~~would~~^{will} be subject to ~~a~~ strict control in the future.

15 Compounds having a high fluorine (F) content as a molecule constituent such as CF_4 , C_2F_6 , etc. have a higher electronegativity of fluorine and thus are chemically very stable. From this nature it is very hard to decompose such fluorine compounds, and ~~it is~~ thus in
20 the current situations that no appropriate processes for such decomposition treatment are not available yet.

JP-B-6-59388 (US Patent No. 5,176,897) discloses a TiO_2 - WO_3 catalyst for hydrolysis of organic halogen compounds. The catalyst contains 0.1 to 20% by
25 weight of W on the basis of TiO_2 (i.e. 92% to 99.96% of Ti

by atom and 8 to 0.04% by atom of W) and has a decomposition rate of 99% at 375°C for a duration of 1,500 hours in treatment of CCl_4 in ppm order. JP-B-6-59388 suggests that organic halogen compounds having a single carbon atom, such as CF_4 , CCl_2F_2 , etc. can be decomposed, but shows no examples of decomposition results of fluorine compounds.

JP-A-7-80303 discloses another $\text{Al}_2\text{O}_3\text{-ZrO}_2\text{-WO}_3$ catalyst for decomposition of fluorine compound gases.

10 The catalyst is directed to combustion-decomposition of CFCs (chlorofluorocarbons) and has a decomposition rate of 98% for a duration of 10 hours in treatment of CFC-115 (C_2ClF_5) by combustion-decomposition reaction at 600°C. The disclosed process needs addition of hydrocarbons such as n-butane, etc. as a combustion aid, resulting in a higher treatment cost. Among organic halogen compounds to be treated, fluorine compounds are less decomposable than chlorine compounds. Furthermore, the more the carbon atoms of organic halogen compound, the less decomposable. Decomposition of compounds consisting only of carbon and fluorine such as C_2F_6 , etc. are much less decomposable than CFC-115, but no examples of decomposition results of such compounds are shown therein.

25 SUMMARY OF THE INVENTION

An object of the present invention is to provide a process for efficient decomposition treatment

of compounds of carbon and fluorine, compounds of carbon,
hydrogen and fluorine, compounds of sulfur and fluorine,
compounds of nitrogen and fluorine and even compounds of
carbon, hydrogen, fluorine and oxygen such as C_2F_6 , CF_4 ,
5 C_3F_8 , C_4F_8 , CHF_3 , SF_6 and NF_3 .

The present invention provides a process for
treating a fluorine compound-containing gas, which
comprises contacting a gas stream containing at least one
fluorine compound selected from the group consisting of
10 compounds of carbon and fluorine, compounds of carbon,
hydrogen and fluorine, compounds of sulfur and fluorine,
compounds of nitrogen and fluorine and compounds of
carbon, hydrogen, oxygen and fluorine with a catalyst
containing at least one of alumina, titania, zirconia and
15 silica in the presence of steam, thereby hydrolyzing the
fluorine compound to ^{Convert}~~convert~~ fluorine of the fluorine
compound to hydrogen fluoride.

BRIEF DESCRIPTION OF THE DRAWINGS

Fig. 1 is a block diagram showing a process for
20 treating a fluorine compound-containing gas according to
one embodiment of the present invention.

Fig. 2 is a graph showing performances of
various catalysts for decomposing a fluorine compound.

Fig. 3 is a graph showing performances of
25 various catalysts for decomposing a fluorine compound.

Fig. 4 is a graph showing performances of
various catalysts for decomposing a fluorine compound.

Fig. 5 is a graph showing performances of various catalysts for decomposing a fluorine compound.

Fig. 6 is a graph showing performances of various catalysts for decomposing a fluorine gas.

5 Fig. 7 is a graph showing performance of catalysts with various composition ratios for decomposing a fluorine gas.

Fig. 8 is a graph showing performance of catalysts with various composition ratios for decomposing
10 a fluorine gas.

Fig. 9 is a graph showing relations between reaction temperature and decomposition rate of various fluorine compounds.

Fig. 10 is a graph showing relations between
15 reaction time and decomposition rate of a fluorine compound.

Fig. 11 is a graph showing relations between reaction temperature and decomposition rate of CHF_3 , CF_4 and C_4H_8 by an Al_2O_3 -ZnO catalyst.

20 Fig. 12 is a graph showing relations between reaction temperature and decomposition rate of SF_6 and C_3F_8 by an Al_2O_3 -NiO catalyst.

Fig. 13 is a graph showing relations between reaction temperature and decomposition rate of C_4F_8 by an
25 Al_2O_3 -NiO-ZnO catalyst.

DESCRIPTION OF THE PREFERRED EMBODIMENTS

As a result of extensive studies on development

of catalysts for decomposition of fluorine compound-
containing gases, the present inventors have found that
catalysts must contain a metallic component capable of
forming an appropriately strong bond with fluorine as the
5 nature of catalysts, and further have found that
catalysts containing a metallic component having a higher
fluoride formation enthalpy show a higher decomposition
activity particularly in case of compounds consisting of
carbon and fluorine, because molecules of such compounds
10 are stable by themselves. Formation of too stable a bond
will gradually lower the decomposition activity of
catalysts, because fluorine compounds are less releasable
from the catalyst surface, whereas too weak a bonding
force will not attain a satisfactory decomposition rate.
15 C_2F_6 , one of gases to be treated according to the present
invention, is a compound of poor reactivity because of a
higher intramolecular force, and it is said that a
temperature of $1,500^{\circ}$ to $2,000^{\circ}C$ is required for combus-
tion of such a gas.

20 As a result of tests on various catalysts, the
present inventors have found that catalysts of alumina
(Al_2O_3), titania (TiO_2), zirconia (ZrO_2), silica (SiO_2), a
mixture of titania and zirconia, a mixture of alumina and
magnesia (MgO), a mixture of alumina and titania, or a
25 mixture of alumina and silica can hydrolyze fluorine
compounds, and further have found that the fluorine
compounds can be decomposed at a lower temperature than
 $800^{\circ}C$ thereby.

Among these catalysts, it has been found that a catalyst based on a mixture of alumina and titania has the highest activity and particularly a catalyst comprising 75 to 98% by weight of alumina and 25 to 2% by weight of titania has a particularly high activity. It can be presumed that the alumina of the catalyst based on a mixture of alumina and titania acts to attract fluorine compounds onto the catalyst, whereas the titania acts to depart the fluorine compounds from the catalyst surface.

10 The present inventors further have found that catalysts based on the mixture of alumina and titania further containing at least one member selected from the group consisting of zirconia, tungsten oxide, silica, tin oxide, ceria, bismuth oxide, nickel oxide and boron oxide
15 can hydrolyze fluorine compounds. It has been found that above all, the catalyst containing zirconia has a higher decomposition activity on fluorine compounds. It has been further found that the content of at least one member selected from the group consisting of zirconia,
20 tungsten oxide, silica, tin oxide, ceria, bismuth oxide, nickel oxide and boron oxide is preferably 0.1 to 10% by weight on the basis of sum total of ^{the} alumina and titania and particularly the content of zirconia is preferably 2 to 10% by weight on the basis of ^{the} sum total of alumina and
25 titania. It seems that these additive members exist in the form of single oxides or composite oxides and contribute to an improvement of decomposition activity on fluorine compounds.

In catalyst preparation, it has been found that it is preferable to use boehmite for alumina raw material and a titanium sulfate solution for a titania raw material. It has been confirmed that there are sulfate ions, SO_4^{2-} , in the catalysts prepared from the titanium sulfate solution and the decomposition activity on fluorine compounds can be improved by the presence of sulfate ions. It has been found that ^{the} addition of sulfuric acid is preferable during the catalyst preparation.

The present inventors further tested catalysts containing other components besides alumina and titania, specifically catalysts containing alumina and one of zinc oxide (ZnO), nickel oxide (NiO), iron oxide, tin oxide (SnO_2), platinum (Pt), cobalt oxide, zirconia (ZrO_2), ceria (CeO_2) and silica (SiO_2). As a result, it has been found that these catalysts can hydrolyze fluorine compounds and particularly catalysts containing zinc oxide or nickel oxide have a higher activity than catalysts based on the mixture of alumina and titania. It has been further found that catalysts comprising alumina and nickel oxide, admixed with sulfuric acid during the catalyst preparation have a higher activity than the catalyst without admixing with sulfuric acid. It has not been confirmed in which forms iron oxide or cobalt oxide of the catalysts containing the iron oxide or the cobalt oxide exists. Probably it seems to exist in the form of Fe_2O_3 or Co_3O_4 .

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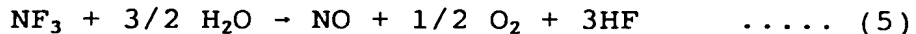
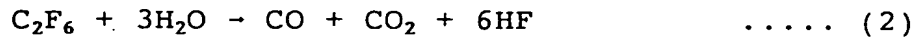
It has been found that the catalysts comprising alumina and one of Zinc oxide, nickel oxide, iron oxide, tin oxide, cobalt oxide, zirconia, ceria and silica as other components preferably contain 50 to 1% by atom of one metallic element of the other components, the balance being aluminum of the alumina, and the content of platinum is preferably 0.1 to 2% by weight on the basis (100% by weight) of alumina. It has been further found that these catalysts can further contain sulfur and the content of sulfur is preferably 0.1 to 20% by weight on the basis of the alumina catalyst.

Fluorine compounds to be treated according to the present invention include, for example, CF_4 , C_2F_6 , C_3F_8 , C_4F_8 , C_5F_8 , CHF_3 , CH_2F_2 , CH_3F , C_2HF_5 , $\text{C}_2\text{H}_2\text{F}_4$, $\text{C}_2\text{H}_3\text{F}_3$, $\text{C}_2\text{H}_4\text{F}_2$, $\text{C}_2\text{H}_5\text{F}$, CH_2OCF_2 , SF_6 , NF_3 , etc., among which CF_4 , C_2F_6 , C_3F_8 , C_4F_8 , CHF_3 , SF_6 and NF_3 are used as etchants for semiconductors and CF_4 , C_2F_6 and NF_3 are used as cleaners for semiconductors.

According to the present invention, all of these fluorine compounds can be hydrolyzed. Hydrolysis temperature depends upon ^{the} kinds of fluorine compounds and catalyst components, and is usually 200° to 800°C , preferably 400° to 800°C . According to the present process fluorine of fluorine compound gases can be converted to hydrogen fluoride.

Hydrolysis of fluorine compounds can proceed typically according to the following reaction equations:





5 Hydrolysis according to reaction equations (2) and (3) can produce CO. The present catalysts also have an ability to oxidize CO, and thus CO can be further oxidized to CO₂ in the presence of oxygen.

 The present invention provides a process for
10 hydrolyzing a fluorine compound-containing gas by a catalyst comprising at least one member selected from the group consisting of alumina, titania, zirconia, silica, a mixture of titania and zirconia, a mixture of alumina and magnesia, a mixture of alumina and titania and a mixture
15 of alumina and silica.

 Furthermore, the present invention provides a process for treating a fluorine-containing gas by a catalyst comprising alumina and titania, further containing 0.1 to 10% by weight, on the basis of alumina and
20 titania, of one of zirconia, tungsten oxide, silica, tin oxide, ceria, bismuth oxide, nickel oxide and boron oxide.

 Still furthermore, the present invention provides a process for treating a fluorine compound-
25 containing gas by a catalyst comprising alumina and at least one member selected from the group consisting of zinc oxide, nickel oxide, iron oxide, tin oxide, cobalt oxide, zirconia, ceria, silica and platinum as other

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components, a ratio of aluminum of alumina to the
metallic element of at least one of ^{the} other components by
atom is 50 to 99 : 50 - 1, and further by the catalyst
further containing 0.1 to 20% by weight of sulfur on the
5 basis of the alumina. These additive components can
contribute to improvement of decomposition activity of
the catalysts on fluorine compounds in the form of single
oxides or composite oxides with aluminum and/or other
additive components.

10 Still furthermore, the present invention
provides a process for converting fluorine in a gas to
hydrogen fluoride, which comprises contacting a gas
stream containing a fluorine compound comprising C_2F_6 with
a catalyst comprising a mixture of alumina and titania
15 and having a weight ratio of alumina to titania ^{of} ~~being~~ 65
to 90 : 35 to 10, thereby hydrolyzing the fluorine
compounds. Still furthermore, the present invention
provides a process for converting fluorine in a gas
stream to hydrogen fluoride, which comprises contacting a
20 gas stream comprising a fluorine compound comprising C_2F_6
with a catalyst comprising a mixture of alumina, titania
and zirconia and having a weight ratio of alumina to
titania being 65 to 90 : 35 to 10 and a weight ratio of
zirconia to ^{the} ~~sum~~ total of alumina and titania being 2 to
25 10 : 98 to 90, thereby hydrolyzing the fluorine compound.

Still furthermore, the present invention
provides a process for converting fluorine in a gas
stream to hydrogen fluoride, which comprises contacting a

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gas stream containing at least one fluorine compound
selected from the group consisting of C_2F_6 , CF_4 , C_4F_8 and
 CHF_3 , with a catalyst comprising a mixture of alumina and
zinc oxide and having an atomic ratio of aluminum to zinc
5 being 90 to 70 : 10 to 30, thereby hydrolyzing the
fluorine compound.

Still furthermore, the present invention
provides a process for converting fluorine in a gas
stream to hydrogen fluoride, which comprises contacting a
10 gas stream containing at least one fluorine compound
selected from the group consisting of C_2F_6 , CF_4 , C_3F_8 , CHF_3 ,
 NF_3 and SF_6 with a catalyst comprising a mixture of
alumina and nickel oxide and having an atomic ratio of
aluminum to nickel of 95 to 60 : 5 to 40, thereby
15 hydrolyzing the fluorine compound.

Still furthermore, the present invention
provides a process for converting fluorine in a gas
stream to hydrogen fluoride, which comprises contacting a
gas stream comprising a fluorine compound comprising C_4F_8
20 with a catalyst comprising a mixture of alumina, nickel
oxide and zinc oxide, thereby hydrolyzing the fluorine
compound.

Still furthermore, the present invention
provides a process for converting fluorine in a gas
25 stream to hydrogen fluoride, which comprises a hydrolysis
step of contacting a gas discharged from a semiconductor-
etching or cleaning step using a gas stream containing at
least one fluorine compound selected from the group

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consisting of compounds of carbon and fluorine, compounds of carbon, hydrogen and fluorine, compounds of sulfur and fluorine, compounds of nitrogen and fluorine and compounds of carbon, hydrogen, oxygen and fluorine, after
5 addition of air and steam to the gas, with a catalyst comprising at least one of alumina, titania, zirconia and silica, thereby hydrolyzing the fluorine compound to convert the fluorine in the gas to hydrogen fluoride, as a poststep to the semiconductor-etching or cleaning step.

10 Still furthermore, the present invention provides a process for treating a fluorine compound-containing gas, which further comprises an alkaline washing step of contacting the gas from the hydrolysis step with an alkaline washing solution, thereby washing
15 the gas as a poststep to the hydrolysis step. As the alkaline washing solution, there can be used conventional ones such as a solution of NaOH, $\text{Ca}(\text{OH})_2$, $\text{Mg}(\text{OH})_2$, CaCO_3 , etc., a slurry of $\text{Ca}(\text{OH})_2$, etc.

In contacting of the gas stream containing a
20 fluorine compound with the catalyst, the concentration of the fluorine compound in the gas stream is preferably 0.1 to 10% by volume, particularly preferably 0.1 to 3% by volume, and the space velocity is preferably 100 to 10,000 h^{-1} , particularly preferably 100 to 3,000 h^{-1} .
25 Space velocity (h^{-1}) is defined by reaction gas flow rate (ml/h)/catalyst volume (ml).

In the hydrolysis of the fluorine compound, it is desirable to add steam as a hydrogen source for

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Bubbling of the decomposition product gas through the alkaline solution or washing with the alkaline solution through a packed column may be used for the neutraliza-

As the raw material for aluminum (Al) for preparing the present catalyst, γ -alumina and a mixture of γ -alumina and δ -alumina can be used besides boehmite. However, it is preferable to use boehmite as a raw material for Al to form an oxide through final firing.

As the third metallic components for silica (Si), magnesium (Mg), zirconium (Zr), etc., their various nitrates, ammonium salts, chlorides, etc. can be used.

The present catalyst can be used as such or upon molding into a granular form, a honeycomb form, etc. by ²an desired molding procedure such as extrusion molding, tableting, tumbling granulation, etc., or as a coating on ceramic or metallic honeycombs or plates.

Only a catalytic reactor for decomposing
fluorine compounds and a facility for neutralizing and

removing acid components in the decomposition product gas are required for an apparatus for carrying out the present process for treating fluorine compound-containing gas.

5 The present invention will be described in detail below, referring to Examples which are not limitative of the present invention.

Fig. 1 shows an example of using the present process for hydrolysis treatment of a fluorine gas in a
10 cleaning step in a plasma CVD apparatus in the semiconductor production process.

The plasma CVD apparatus is an apparatus for vapor depositing a SiO_2 film on a semiconductor wafer surface. Since the SiO_2 film tends to deposit on the
15 entire interior surfaces of the apparatus, and thus it is necessary to remove SiO_2 depositions from unwanted surfaces. To clean the unwanted surface to remove SiO_2 therefrom, gases containing fluorine compounds such as C_2F_6 , CF_4 , NF_3 , etc. are used as a cleaning gas. Cleaning
20 gas 1 containing these fluorine compounds is led to a CVD chamber to remove SiO_2 under plasma excitation. Then, the chamber is flushed with a N_2 gas 2, thereby diluting the cleaning gas to a desired lower fluorine compound concentration, and the diluted cleaning gas is discharged
25 from the chamber. The discharged gas is admixed with air 3 to further lower the fluorine compound concentration by dilution with air 3 and the air-diluted discharged gas is further admixed with steam 4 and the resulting reaction

gas 5 is led to a decomposition step, where the reaction gas is brought into contact with a catalyst at a desired space velocity (h^{-1}), which is defined by reaction gas flow rate (ml/h)/catalyst volume (ml) and at a desired temperature. In that case, the reaction gas may be heated or the catalyst may be heated by an electric oven, etc. The resulting decomposition gas 6 is led to an exhaust gas washing step, where the decomposition gas 6 is sprayed with an aqueous alkaline solution to remove acid components from the decomposition gas 6 and the resulting exhaust gas 7 freed from the acid components is discharged to the system outside.

CF_4 , C_2F_6 and NF_3 can be used as etchants for semiconductors, etc., and CHF_3 , C_3F_8 , SF_6 and C_4F_8 can be also used as etchants besides the above-mentioned fluorine compounds. These etchants can be treated and decomposed in the same manner as in Fig. 1 except that the cleaning step of Fig. 1 is only replaced with an etching step.

Activities or performances of various catalysts for composing fluorine compounds were investigated, and results thereof will be described below:

Example 1

A C_2F_6 gas having a purity of 99% or more was diluted with air, and further admixed with steam to prepare a reaction gas. Steam for the admixture was prepared by feeding pure water into a reactor tube from

the top at a flow rate of 0.11 ml/min. by a microtube pump and gasified. The reaction gas had a C_2F_6 concentration of about 0.5%. Then, the reaction gas was brought into contact with various catalysts heated to 700°C in a reactor tube at a space velocity of 3,000 h⁻¹. Heating of the catalyst was carried out by heating the reactor tube in an electric oven.

Ans E2 → *the reactor*
Reactor tube was an Inconel reactor tube having an inner diameter of 19 mm, where a catalyst bed was fixed at the center of the reactor tube and had an Inconel thermowell for a thermo couple, 3 mm in outer diameter, inside the catalyst bed. Decomposition product gas discharged from the catalyst bed was bubbled through an aqueous sodium chloride solution *and* then discharged as an exhaust gas. *The* C_2F_6 decomposition rate was calculated by the following equation by determining concentration of C_2F_6 in the reaction gas at the inlet to the reactor tube and concentration of C_2F_6 in the decomposition gas at the outlet from the alkaline washing step by FID (flame ionization detector) gas chromatography and TCD (thermal conductivity detector) gas chromatography:

$$\text{Decomposition rate} = 1 - \frac{\text{Concentration of discharged fluorine compound}}{\text{Concentration of fed fluorine compound}} \times 100 (\%)$$

Catalyst 1 : Al_2O_3

Granular alumina (NKHD-24, *the* trademark of a product commercially available from Sumitomo Chemical

Co., Ltd., Japan) was pulverized, sieved to obtain a fraction of 0.5 - 1 mm grain sizes, followed by drying at 120°C for 2 hours and firing (or calcining) at 700°C for 2 hours.

5 Catalyst 2 : TiO_2

Granular titania (CS-200-24, ^{the} trademark of a product commercially available from Sakai Chemical Industry Co., Ltd., Japan) was pulverized, sieved to obtain a fraction of 0.5 - 1 mm grain sizes, followed by
10 drying at 120°C for 2 hours and firing at 700°C for 2 hours.

Catalyst 3 : ZrO_2

200 g of zirconyl nitrate was dried at 120°C for 2 hours and fired at 700°C for 2 hours. The result-
15 ing powders were placed in a mold and compression molded under a pressure of 500 kgf/cm². The molded product was pulverized and sieved to obtain zirconia grains having grain sizes of 0.5 - 1 mm.

Catalyst 4 : SiO_2

20 Granular silica (CARIACT-10, ^{the} trademark of a product commercially available from Fuji Silysia Co., Ltd., Japan) was pulverized and sieved to obtain a fraction of 0.5 - 1 mm grain sizes, followed by drying at 120°C for 2 hours and firing at 700°C for 2 hours.

25 Catalyst 5 : $\text{TiO}_2\text{-ZrO}_2$

Granular titania (CS-200-24) was pulverized to grain sizes of 0.5 mm and under. 100 g of the resulting powders was admixed with 78.3 g of zirconyl nitrate and

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5 under a pressure of 500 kgf/cm². The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Ti : Zr = 81 : 19 and in a weight ratio of TiO₂ : ZrO₂ = 73.5 : 26.5.

Granular alumina (NKHD-24) was pulverized to grain sizes of 0.5 mm and under. 100 g of the resulting powders were admixed with 56.4 g of magnesium nitrate and kneaded while adding pure water thereto. After the kneading, the kneaded mixture was dried at 120°C for 2 hours and fired at 700°C for 2 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm². The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Mg = 90 : 10 and in a weight ratio of Al₂O₃ : MgO = 91.9 : 8.1.

Granular alumina (NKHD-24) was pulverized to grain sizes of 0.5 mm and under. 100 g of the resulting powders were admixed with 17.4 g of dried powders of a metatitanic acid slurry and kneaded while adding pure water thereto. After the kneading, the kneaded mixture

was dried at 120°C for 2 hours and fired at 700°C for 2 hours. The resulting powders were placed in a mold and compression molded under a pressure of 500 kgf/cm². The molded product was pulverized and sieved to obtain grains
5 having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti = 90 : 10 and in a weight ratio of Al₂O₃ : TiO₂ = 85.2 : 14.8.

Catalyst 8 : Al₂O₃-SiO₂

10 Granular alumina (NKHD-24) was pulverized to grain sizes of 0.5 mm and under. 100 g of the resulting powders were admixed with 13.2 g of dried powders of SiO₂ sol and kneaded while adding pure water thereto. After the kneading, the kneaded mixture was dried at 120°C for
15 2 hours and fired at 700°C for 2 hours. The resulting powders were placed in a mold and compression molded under a pressure of 500 mgf/cm². The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for
20 catalyst was in an atomic ratio of Al : Si = 90 : 10 and in a weight ratio of Al₂O₃ : SiO₂ = 88.3 : 11.7.

Test results of the above-mentioned catalysts 1 to 8 are shown in Fig. 2, from which it is evident that the Al₂O₃-TiO₂ catalyst is preferable as a hydrolysis
25 catalyst for a C₂F₆ gas.

Example 2

In this Example, influences of changes in

composition ratios of alumina to titania in $\text{Al}_2\text{O}_3\text{-TiO}_2$ catalysts upon C_2F_6 decomposition rate were investigated under the same test procedure and conditions as in Example 1. The results are shown in Fig. 4.

5 Catalyst 19 : Al_2O_3

Boehmite powders (PURAL SB, trademark of a product commercially available from Condea Co., Ltd.) were dried at 120°C for 2 hours. 200 g of the resulting dried powders were fired at 300°C for 0.5 hours and
10 further fired at an elevated temperature of 700°C for 2 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm^2 . The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm, and tested. It was
15 found that boehmite powders used as an alumina raw material had a higher catalytic activity than granular alumina.

Catalyst 20 : $\text{Al}_2\text{O}_3\text{-TiO}_2$

Boehmite powders (PURAL SB) were dried at 120°C
20 for one hour. 200 g of the resulting dried powders were kneaded with 248.4 g of an aqueous 30% titanium sulfate solution, while adding about 200 g of pure water thereto. After the kneading, the kneaded mixture was dried at 250°C - 300°C for about 5 hours and then fired at 700°C for 2
25 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm^2 . The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The result-

ing grain composition for catalyst was in an atomic ratio of $Al : Ti = 90 : 10$ and in a weight ratio of $Al_2O_3 : TiO_2 = 85.65 : 14.35$.

Catalyst 21 : AlO_3-TiO_2

5 Boehmite powders (PURAL SB) were dried at $120^\circ C$ for one hour. 200 g of the resulting dried powders were kneaded with about 100 g of an aqueous solution containing 78.6 g of 30% titania sol in pure water. After the kneading, the kneaded mixture was dried at
10 $120^\circ C$ for about 2 hours and then fired at $700^\circ C$ for 2 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm^2 . The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The result-
15 ing grain composition for catalyst was in an atomic ratio of $Al : Ti = 91 : 9$ and in a weight ratio of $Al_2O_3 : TiO_2 = 86.25 : 13.75$.

It was found that the catalyst prepared from the titanium sulfate solution as a titanium raw material
20 had the highest catalytic activity, probably because of the presence of sulfate ions SO_4^{2-} in the catalyst.

Example 3

In this Example, influences of changes in composition ratios of Al_2O_3 to TiO_2 in $Al_2O_3-TiO_2$ catalysts
25 upon C_2F_6 decomposition rate were investigated under the same procedure and conditions as in Example 1.

Catalyst 22 : $\text{Al}_2\text{O}_3\text{-TiO}_2$

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 100 g of the resulting dried powders were kneaded with 82.4 g of an aqueous 30% titanium sulfate solution while adding about 120 g of pure water thereto. After the kneading, the kneaded mixture was dried at 250°C - 300°C for about 5 hours and then fired at 700°C for 2 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm^2 . The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of $\text{Al} : \text{Ti} = 93 : 7$ and in a weight ratio of $\text{Al}_2\text{O}_3 : \text{TiO}_2 = 90.0 : 10.0$.

15 Catalyst 23 : $\text{Al}_2\text{O}_3\text{-TiO}_2$

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 100 g of the resulting dried powders were kneaded with 174.4 g of an aqueous 30% titanium sulfate solution while adding about 70 g of pure water thereto. After the kneading, the kneaded mixture was dried at 250°C - 300°C for about 5 hours and then fired at 700°C for 2 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm^2 . The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of $\text{Al} : \text{Ti} = 87 : 13$ and in a weight ratio of $\text{Al}_2\text{O}_3 : \text{TiO}_2 = 80.9 : 19.1$.

Catalyst 24 : $\text{Al}_2\text{O}_3\text{-TiO}_2$

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 100 g of the resulting dried powders were kneaded with 392 g of an aqueous 30% titanium sulfate solution while adding the latter to the former. After the kneading, the kneaded mixture was dried at $250^\circ - 300^\circ\text{C}$ for about 5 hours and then fired at 700°C for 2 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm^2 . The molded product was pulverized and sieved to obtain grains having grain sizes of $0.5 - 1\text{ mm}$ and tested. The resulting grain composition for catalyst was in an atomic ratio of $\text{Al} : \text{Ti} = 75 : 25$ and in a weight ratio of $\text{Al}_2\text{O}_3 : \text{TiO}_2 = 65.4 : 34.6$.

Activities of catalysts 19, 20 and 22-24 are shown in Fig. 5, from which is evident that the highest C_2F_6 decomposition rate can be obtained at an alumina content of about 85% by weight.

Example 4

In this Example, an influence of sulfuric acid during the preparation of the $\text{Al}_2\text{O}_3\text{-TiO}_2$ catalyst upon the C_2F_6 decomposition rate was investigated.

Catalyst 25 : $\text{Al}_2\text{O}_3\text{-TiO}_2$

Boehmite powders (PURAL SB) was dried at 120°C for one hour. 150 g of the resulting dried powders were kneaded with 58.5 g of 30% titania sol (CS-N, ^{the} trademark of a product commercially available from Ishihara Sangyo

Kaisha, Ltd., Japan) and an aqueous solution prepared by diluting 44.8 g of 97% sulfuric acid with 250 ml of pure water. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 5 hours and then fired at 700°C for 2 hours. The resulting powders were placed into a mold and compression molded under a pressure of 500 kgf/cm². The molded product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti = 91 : 9 and in a weight ratio of Al₂O₃ : TiO₂ = 86.3 : 13.7.

Sulfate ions were present in the catalyst. Test conditions were the same as in Example 1, except that the space velocity was changed to 1,000 h⁻¹. The test results revealed that a C₂F₆ decomposition rate of 80% was obtained at a reaction temperature of 650°C.

Example 5

In this Example, C₂F₆ decomposition rates were investigated by adding various components to the Al₂O₃-TiO₂ catalysts. The catalysts were prepared as follows, but test procedure and conditions were the same as in Example 1.

Catalyst 9 : Al₂O₃-TiO₂

Granular alumina (NKHD-24) was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm, followed by drying at 120°C for 2 hours. Then, the dried grains were impregnated with 176 g of an aqueous 30%

titanium sulfate solution. After the impregnation, the grains were dried at 250° - 300°C for about 5 hours and then fired at 700°C for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al :

5 Ti = 90 : 10 and in a weight ratio of Al_2O_3 : TiO_2 = 85.1 : 14.9. The catalyst thus prepared was designated as catalyst A.

Catalyst 10 : Al_2O_3 - TiO_2 - ZrO_2

50 g of Catalyst A grains were impregnated with
10 an aqueous solution of 6.7 g of zirconyl nitrate dihydrate in 38.4 g of H_2O . After the impregnation, the grains were dried at 120°C for 2 hours and then fired at 700°C for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti : Zr = 90 : 10
15 : 0.025 and in a weight ratio of Al_2O_3 : TiO_2 : ZrO_2 = 80.2 : 14.0 : 5.8.

Catalyst 11 : Al_2O_3 - TiO_2 - WO_3

50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 6.5 g of ammonium
20 paratungstate in H_2O . After the impregnation, the grains were dried at 120°C for 2 hours and then fired at 700°C for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti : W = 90 : 10 : 0.025 and in a weight ratio of Al_2O_3 : TiO_2 : WO_3 = 76.6
25 : 13.4 : 10.0.

Catalyst 12 : Al_2O_3 - TiO_2 - SiO_2

50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 7.5 g of 20 wt.% silica

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sol in H_2O . After the impregnation, the grains were dried at $120^\circ C$ for 2 hours and then fired at $700^\circ C$ for 2 hours.

The resulting grain composition for catalyst was in an atomic ratio of 90 : 10 : 0.025 and in a weight ratio of

5 Al_2O_3 : TiO_2 : SiO_2 = 82.6 : 14.5 : 2.9.

Catalyst 13 : Al_2O_3 - TiO_2 - SnO_2

50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 5.6 g of tin chloride dihydrate in H_2O . After the impregnation, the grains were

10 dried at $120^\circ C$ for 2 hours and then fired at $700^\circ C$ for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti : Sn = 90 : 10 : 0.025 and in a weight ratio of Al_2O_3 : TiO_2 : SnO_2 = 79.1 : 13.9 : 7.0.

15 Catalyst 14 : Al_2O_3 - TiO_2 - CeO_2

50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 10.9 g of cerium nitrate hexahydrate in H_2O . After the impregnation, the grains were dried at $120^\circ C$ for 2 hours and then fired at $700^\circ C$

20 for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti : Ce = 90 : 10 : 0.025 and in a weight ratio of Al_2O_3 : TiO_2 : CeO_2 = 78.4 : 13.7 : 7.9.

Catalyst 15 : Al_2O_3 - TiO_2 - MnO_2

25 50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 7.2 g of manganese nitrate hexahydrate in H_2O . After the impregnation, the grains were dried at $120^\circ C$ for 2 hours and then fired at

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700°C for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti : Mn = 90 : 10 : 0.025 and in a weight ratio of Al₂O₃ : TiO₂ : MnO₂ = 81.6 : 14.3 : 4.1.

5 Catalyst 16 : Al₂O₃-TiO₂-Bi₂O₃

50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 12.13 g of ~~bismuth~~ ^{bismuth} nitrate hexahydrate in H₂O. After the impregnation, the grains were dried at 120°C for 2 hours and then fired at 700°C for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti : Bi = 90 : 10 : 0.025 and in a weight ratio of Al₂O₃ : TiO₂ : Bi₂O₃ = 85.1 : 14.8 : 1.1.

Catalyst 17 : Al₂O₃-TiO₂-NiO

15 50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 7.3 g of nickel nitrate hexahydrate in H₂O. After the impregnation, the grains were dried at 120°C for 2 hours and then fired at 700°C for 2 hours. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti : Ni = 90 : 10 : 0.025 and in a weight ratio of Al₂O₃ : TiO₂ : NiO = 82.0 : 14.4 : 3.6.

Catalyst 18 : Al₂O₃-TiO₂-BO₄

25 50 g of Catalyst A grains were impregnated with 38.4 g of an aqueous solution of 1.36 g of ammonium borate octahydrate in H₂O. After the impregnation, the grains were dried at 120°C for 2 hours and then fired at 700°C for 2 hours. The resulting grain composition for

catalyst was in an atomic ratio of Al : Ti : B = 90 : 10 : 0.005 and in a weight ratio of Al_2O_3 : TiO_2 : BO_3 = 85.65 : 14.827 : 0.008.

It was found from Fig. 3 that the Al_2O_3 - TiO_2 - ZrO_2 catalyst had the highest activity.

Example 6

In this Example, various catalysts containing alumina as one member were investigated for C_2F_6 decomposition rates under the following conditions:

10 A C_2F_6 gas having a purity of 99% or more was diluted with air, and the diluted gas was further admixed with steam to prepare a reaction gas. Steam was prepared by feeding pure water to a reactor tube from the top at a flow rate of about 0.2 ml/min. by a microtube pump to
15 gasify the pure water. The reaction gas had a C_2F_6 concentration of about 0.5%, and was brought into contact with a catalyst heated to 700°C by external heating of the reactor tube in an electric oven at a space velocity of 2,000 h^{-1} .

20 The reactor tube was an Inconel reactor tube having an inner diameter of 32 mm and had a catalyst bed fixed at the center of the reactor tube. An Inconel thermowell for a thermocouple, 3 mm in diameter, was inserted into the catalyst bed. Decomposition product
25 gas from the catalyst bed was bubbled through an aqueous calcium fluoride solution and discharged to the system outside.

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The following catalysts were prepared for the test under the foregoing conditions:

Catalyst 26

Boehmite powders (PURAL SB) were dried at 120°C
5 for one hour. 200 g of the resulting dried powders zinc
nitrate hexahydrate and the mixture was kneaded. After
the kneading, the kneaded mixture was dried at 250° -
300°C for about 2 hours and then fired at 700°C for 2
hours. The fired product was pulverized and sieved to
10 obtain grains having grain sizes of 0.5 - 1 mm and
tested. The resulting grain composition for catalyst was
in an atomic ratio of Al : Zn = 91 : 9 and in a weight
ratio of Al₂O₃ : ZnO = 86.4 : 13.6.

Catalyst 27

15 Boehmite powders (PURAL SB) were dried at 120°C
for one hour. 200 g of the resulting dried powders were
admixed with an aqueous solution of 50.99 g of nickel
sulfate hexahydrate and the mixture was kneaded. After
the kneading, the kneaded mixture was dried at 250° -
20 300°C for about 2 hours and then fired at 700°C for 2
hours. The fired product was pulverized and sieved to
obtain grains having grain sizes of 0.5 - 1 mm. The
resulting grain composition for catalyst was in an atomic
ratio of Al : Ni = 91 : 9 and in a weight ratio of Al₂O₃ :
25 NiO = 87.3 : 12.7.

Catalyst 28

Boehmite ^{powders}~~powder~~ (PURAL SB) were dried at 120°C
for one hour. 300 g of the resulting dried powders were

admixed with an aqueous solution of 125.04 g of nickel nitrate hexahydrate and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Ni = 91 : 9 and in a weight ratio of Al₂O₃ : NiO = 87.3 : 12.7.

10 Catalyst 29

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 300 g of the resulting dried powders were kneaded with 354.4 g of an aqueous 30% titanium sulfate solution while adding about 300 g of pure water thereto. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 5 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of Al : Ti = 91 : 9 and in a weight ratio of Al₂O₃ : TiO₂ = 86.6 : 13.4.

Catalyst 30

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution of 115.95 g of iron nitrate nonahydrate and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2

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hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm, and tested. The resulting grain composition was in an atomic ratio of Al : Fe = 91 : 9.

5 Catalyst 31

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution of 95.43 g of tin chloride hydrate and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then dried at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of Al : Sn = 91 : 9 and in a weight ratio of Al₂O₃ : SnO₂ = 77.4 : 22.6.

Catalyst 32

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution prepared by diluting 22.2 g of a dinitrodiamino Pt(II) nitric acid solution (Pt concentration: 4.5 wt.%) with 200 ml of pure water, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in a weight ratio of

Al_2O_3 : Pt = 100 : 0.68.

Catalyst 33

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 300 g of the resulting dried powders were
5 admixed with an aqueous solution of 125.87 g of cobalt nitrate hexahydrate, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to
10 obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition was in an atomic ratio of Al : Co = 91 : 9.

Catalyst 34

Boehmite powder (PURAL SB) were dried at 120°C
15 for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution of 76.70 g of zirconyl nitrate dihydrate, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2
20 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of Al : Zr = 91 : 9 and in a weight ratio of Al_2O_3 : ZrO_2 = 80.7 : 19.3.

25 Catalyst 35

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution of 124.62 g of cerium

nitrate hexahydrate, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of Al : Ce = 91 : 9 and in a weight ratio of Al₂O₃ : CeO₂ = 75.0 : 25.0.

Catalyst 36

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 300 g of the resulting dried powders were admixed with an aqueous solution of 129.19 g of 20 wt.% silica sol, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm and tested. The resulting grain composition for catalyst was in an atomic ratio of Al : Si = 91 : 9 and in a weight ratio of Al₂O₃ : SiO₂ = 89.6 : 10.4.

(Insert) ²⁴ Test results of the foregoing catalysts 19 and 26 - 36 at a reaction temperature of 700°C are shown in Fig. 6, C₂F₆ decomposition activity is highest with the Al₂O₃-ZnO₂ catalyst and is lowered in the order of the Al₂O₃-NiO catalyst, and the Al₂O₃-TiO₂ catalyst. The highest activity of catalyst 26 seems to be due to the effect of S.

Example 7

In this Example, changes in the activity of Al_2O_3 -NiO catalyst 28 were investigated by changing atomic ratios of Al : Ni. Test procedure and conditions were
5 the same as in Example 6 except that the C_2F_6 concentration was changed to 2% and the feed rate of pure water to 0.4 ml/min.

Catalyst 28-1

Boehmite powders (PURAL SB) were dried at 120°C
10 for one hour. 200 g of the resulting dried powder were admixed with an aqueous solution of 8.52 g of nickel nitrate hexahydrate, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2
15 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Ni = 99 : 1 and in a weight ratio of Al_2O_3 : NiO = 98.5 : 1.5.

20 Catalyst 28-2

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 300 g of the resulting powders were admixed with an aqueous solution of 66.59 g of nickel nitrate hexahydrate, and the mixture was kneaded. After
25 the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The

Catalyst 28-3

Catalyst 28-4

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution of 361.16 g of nickel nitrate hexahydrate, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Ni = 70 : 30 and in a weight ratio of Al₂O₃ : NiO = 61.4 : 38.6.

Catalyst 28-5

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with 562.1 g of nickel nitrate hexahydrate, and the mixture was kneaded while adding water thereto. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Ni = 60 : 40 and in a weight ratio of Al₂O₃ : NiO = 50.6 : 49.4.

The C₂F₆ decomposition rate 6 hours after the start of test is shown in Fig. 7. It was found that the Ni content of Al₂O₃-NiO catalysts is in a range of 5 to 40 atom%, preferably 20 to 30 atom%.

Example 8

In this Example, changes in the activity of Al₂O₃-ZnO catalyst 26 ^{were} ~~was~~ investigated by changing atomic ratios of Al : Zn. Test procedure and conditions were the same as in Example 6 except that the C₂F₆ concentration was changed to 2% and the feed rate of pure water to 0.4 ml/min.

Catalyst 26-1

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution of 215.68 g of zinc

nitrate hexahydrate and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to
5 obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Zn = 80 : 20 and in a weight ratio of Al₂O₃ : ZnO = 71.5 : 28.5.

Catalyst 26-2

10 Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with 369.48 g of zinc nitrate hexahydrate and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and
15 fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic ratio of Al : Zn = 70 : 30 and in a weight ratio of Al₂O₃ : ZnO = 59.4 : 40.6.

20 Catalyst 26-3

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 126.65 g of the resulting dried powders were admixed with an aqueous solution of 96.39 g of zinc nitrate hexahydrate, and the mixture was kneaded. After
25 the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The

resulting grain composition for catalyst was in an atomic ratio of Al : Zn = 85 : 15 and in a weight ratio of Al_2O_3 : ZnO = 78.0 : 22.0.

The C_2F_6 decomposition rate 6 hours after the start of test is shown in Fig. 8. It was found that the Zn content of Al_2O_3 -ZnO catalysts is in a range of 10 to 30 atom%, preferably 10 to 15 atom%.

Example 9

In this Example, decomposition of CF_4 and CHF_3 was carried out with a Al_2O_3 -NiO catalyst 28-3 under the same test procedure and conditions as in Example 6, except that the space velocity was changed to $1,000 \text{ h}^{-1}$ and the fluorine compound was diluted with nitrogen in place of air. Test results at various reaction temperatures are shown in Fig. 9. It was found that the decomposition activity of Al_2O_3 -NiO catalyst upon CF_4 gas and CHF_3 gas is higher than upon C_2F_6 gas and the Al_2O_3 -NiO catalyst is a preferable hydrolysis catalyst for CF_4 or CHF_3 . Furthermore, it was found that a preferable reaction temperature is $600^\circ - 700^\circ\text{C}$ for the decomposition of CF_4 and CHF_3 , and $650^\circ - 700^\circ\text{C}$ for the decomposition of C_2F_6 . The higher the reaction temperature, the higher the decomposition rate. However, substantially 100% decomposition rate can be obtained at 700°C , and thus a higher reaction temperature than 700°C will be meaningless, and a reaction temperature must be as high

as 800°C.

Example 10

In this Example, influences of steam upon C_2F_6 decomposition were investigated under the same test conditions as in Example 6 except that the space velocity was changed 1,000 h^{-1} . Al_2O_3 -NiO catalyst 28-3 was used at a reaction temperature of 700°C while supplying steam for 2 hours from the start of test, then interrupting supply of steam for 3 hours, and then starting to supply steam again. Test results are shown in Fig. 10. It was found that during the supply of steam the C_2F_6 reaction rate was elevated due to the occurrence of C_2F_6 hydrolysis.

Example 11

In this Example, decomposition of SF_6 was investigated with Al_2O_3 -NiO catalyst 28-3 under the same test conditions as in Example 6 except that a SF_6 gas having a purity of 99% or more was used, the space velocity was changed to 1,000 h^{-1} and the SF_6 gas was diluted with nitrogen ^{instead} ~~in~~ of air. The reaction temperature was 700°C. Concentration of SF_6 in the reaction gas at the inlet to the reactor tube and concentration of SF_6 in the decomposition gas at the outlet from the alkaline washing step were determined by TCD gas chromatography and the decomposition rate was calculated by the following equation. It was found that the decomposition rate was 99% or more.

$$\text{Decomposition rate} = 1 - \frac{\text{Concentration of discharged SF}_6}{\text{Concentration of fed SF}_6} \times 100 (\%)$$

Example 12

In this Example, decomposition of NF₃ was investigated with Al₂O₃-NiO catalyst 28-3 under the same test conditions as in Example 11 except that a NF₃ gas having a purity of 99% or more was used. Reaction temperature was 700°C. Concentration of NF₃ in the reaction gas at the inlet to the reactor tube and concentration of NF₃ in the decomposition gas at the outlet from the alkaline washing step were determined by TCD gas chromatography and the decomposition rate was calculated according to the following equation. It was found that the decomposition rate was 99% or more. It was found preferable to carry out the decomposition of the NF₃ gas with the Al₂O₃-NiO catalyst at a temperature of 600° - 800°C.

$$\text{Decomposition rate} = \frac{\text{Concentration of discharged NF}_3}{\text{Concentration of fed NF}_3} \times 100 (\%)$$

Example 13

In this Example, activity of Al₂O₃-ZnO catalyst upon hydrolysis of a CF₄ gas, a C₄F₈ gas and a CHF₃ gas was investigated. Decomposition of a CF₄ gas was carried out in the following manner:

1. The first step is to identify the problem or goal.
 2. The second step is to gather information and resources.
 3. The third step is to develop a plan or strategy.
 4. The fourth step is to implement the plan.
 5. The fifth step is to evaluate the results and make adjustments.

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C₄F₈ as fed and decomposed in the same manner as above. Decomposition rates of CHF₃ and C₄F₈ were determined by FID gas chromatography, whereby it was found that the Al₂O₃-ZnO catalyst had a higher activity upon the CF₄ gas, 5 the C₄F₈ gas and the CHF₃ gas. It was also found that a higher decomposition rate can be obtained preferably at a reaction temperature of 650°C or higher for the hydrolysis of the C₄F₈ gas and even at a reactor temperature of 600°C or lower for the hydrolysis of the CHF₃ gas 10 or the CF₄ gas.

Example 14

In this Example, the decomposition activity of as Al₂O₃-NiO catalyst upon a C₃F₈ gas, a C₄F₈ gas and a SF₆ gas was investigated in the same manner as in Example 13. 15 The concentration of C₄F₈ in the reaction gas after decomposition of C₄F₈ was 0.1% by volume. The Al₂O₃-NiO catalyst used for the test was prepared in the following manner:

Boehmite powders (PURAL SB) were dried at 120°C 20 for one hour. 200 g of the resulting dried powders were admixed with an aqueous solution of 210.82 g of nickel nitrate hexahydrate, and the mixture was kneaded. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours, and then fired at 700°C for 2 25 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting grain composition for catalyst was in an atomic

ratio of Al : Ni = 80 : 20 and in a weight ratio of Al_2O_3 : NiO = 73.2 : 26.8.

Fig. 12 shows decomposition rates at various reaction temperatures, where the decomposition rate of C_3F_8 and C_4F_8 was determined by FID gas chromatography and that of SF_6 by TCD gas chromatography. It was found from the test results that the Al_2O_3 -NiO catalyst had a higher activity upon the hydrolysis of the SF_6 gas, C_3F_8 gas and the C_4F_8 gas, and the reaction temperature was preferably 500°C or higher for the hydrolysis of the SF_6 gas and preferably 700°C or higher for the hydrolysis of the C_3F_8 gas. In the case of C_4F_8 gas, the reaction temperature for the hydrolysis was preferably 650°C or higher.

Example 15

In this Example, decomposition activity of an Al_2O_3 -NiO-ZnO catalyst upon C_4F_8 was investigated in the same manner as in Example 13. The Al_2O_3 -NiO-ZnO catalyst used for the test was prepared in the following manner:

Boehmite powders (PURAL SB) were dried at 120°C for one hour. 200 g of the resulting dried powders were admixed with 210.82 g of nickel nitrate hexahydrate and 152.31 g of zinc nitrate hexahydrate, and the mixture was kneaded while adding pure water thereto. After the kneading, the kneaded mixture was dried at 250° - 300°C for about 2 hours and then fired at 700°C for 2 hours. The fired product was pulverized and sieved to obtain grains having grain sizes of 0.5 - 1 mm. The resulting

Fig. 13 shows decomposition rates at various
5 reaction temperatures, where the decomposition rate of
 C_4F_8 was determined by FID gas chromatography.

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5 reaction temperatures, where the decomposition rate of
 C_4F_8 was determined by FID gas chromatography.